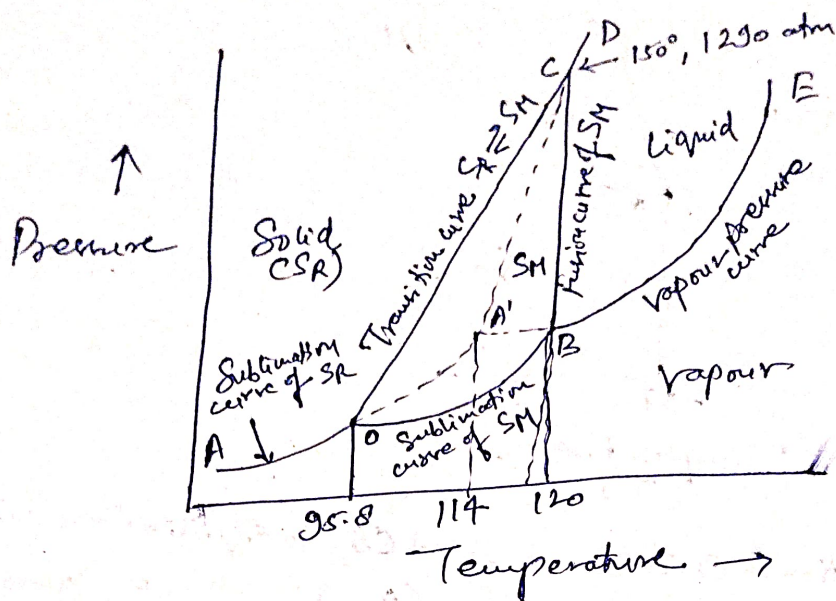


one pure system →



AO → Sublimation curve of rhombic Sulphur → gives vapour pressure at different temperatures.

$(SR \rightleftharpoons SV) \rightarrow$ univariant

O is the transition temperature (95.6°C) at which rhombic Sulphur changes into monoclinic Sulphur.

O is the triple point at which three phases, two solids and the vapour ($SR-SM-SV$) coexist in equilibrium. This is a non-variant point.

O-B is the sublimation curve of monoclinic Sulphur. It gives vapour pressure of monoclinic Sulphur at different temperatures. No. of phases is 2 → univariant. B is the m.p.t (120°C) of SM. This is another triple point at which three phases SM, SL & SV are in equilibrium.

The curve BE is the vapour pressure of liq. Sulphur. $(SL \rightleftharpoons SV) \rightarrow$ univariant.